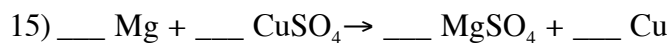
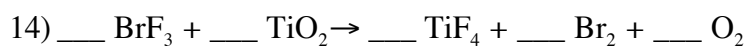
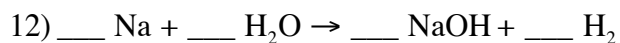
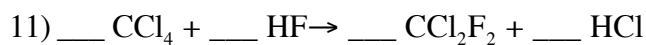
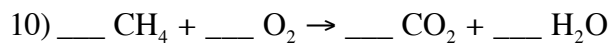
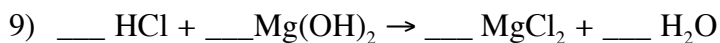
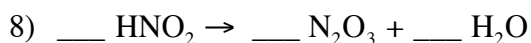
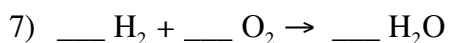
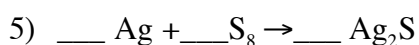
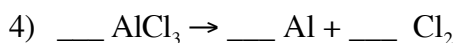
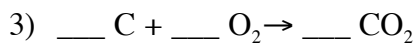
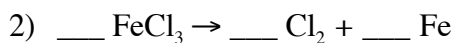


Name _____ Period _____

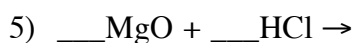
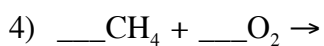
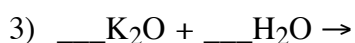
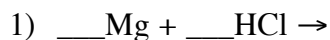
**Chemistry Practice Test
Reactions and Equations**

Form P

Part I: Balance the following equations in the spaces provided. There is no need to include a coefficient of one. Then on the right identify what class of reaction it is.



Part II: Predict the products of the following reactions and balance the equations



Part III: Terms and Symbols Define or identify how you represent each of the following things in a chemical reaction:

Solid

Gas

Liquid

Aqueous

Precipitate

Yields

Produces

Heat

Exothermic

Endothermic

Coefficient

Given the reaction: $\text{HCl} + \text{Mg}(\text{OH})_2 \rightarrow \text{MgCl}_2 + \text{H}_2\text{O}$

Name the Reactants

Name the Products

Part IV: Write and balance the equations represented by the words below.

1) Sodium reacts with water to form sodium hydroxide and hydrogen gas.

2) The combustion of propane (C_3H_8) to make carbon dioxide and water.

3) Hydrochloric acid reacts with sodium hydroxide to make water and sodium chloride.

4) Calcium carbonate decomposes to make carbon dioxide and calcium oxide.

5) Calcium carbonate and hydrochloric acid to produce carbon dioxide, water, and calcium chloride
