

Name _____ Period _____

Honors Chemistry
Matter and Nomenclature Practice Test

Form P

Part I: Fill in formulas for the following combination of ions.

	F ⁻	S ²⁻	SO ₄ ²⁻	PO ₃ ³⁻
Cs ⁺				
Ca ²⁺				
B ³⁺				
K ⁺				
NH ₄ ⁺				

Part II: Write the names of the following Ionic compounds:

A) BaCl₂ _____

B) KCl _____

C) Na₂SO₄ _____

D) K₃PO₄ _____

E) Cu(NO₃)₂ _____

Write the names of the following Molecular Compounds

A) SF₄ _____

B) CO₂ _____

C) N₂O₄ _____

D) H₂O _____

E) P₂O₅ _____

Part III: Write the formula for the compound that is listed:

1. dichlorine monoxide 1. _____

2. nitrogen trifluoride 2. _____

3. chlorine dioxide 3. _____

4. sulfur hexafluoride 4. _____

5. nitrogen dioxide 5. _____

- | | |
|-------------------------|-----------|
| 6. copper (I) fluoride | 6. _____ |
| 7. copper (II) fluoride | 7. _____ |
| 8. ammonium acetate | 8. _____ |
| 9. potassium sulfate | 9. _____ |
| 10. barium chloride | 10. _____ |

Part IV: Define the following terms.

Matter

Ion

Chemical Change

Solution

Physical Change

Element

Atom

Molecule

Compound

Anion

Cation

Monoatomic

Diatomic

Polyatomic

Inorganic

Ionic