

Name \_\_\_\_\_

Period \_\_\_\_\_

### Chemical Formulas and Nomenclature #3

#### Part I: Simple Binary Compounds

Name the following binary compounds:

- |                      |       |                       |       |
|----------------------|-------|-----------------------|-------|
| 1) NaBr              | _____ | 6) AlBr <sub>3</sub>  | _____ |
| 2) KCl               | _____ | 7) MgBr <sub>2</sub>  | _____ |
| 3) Rb <sub>2</sub> O | _____ | 8) BaF <sub>2</sub>   | _____ |
| 4) CaF <sub>2</sub>  | _____ | 9) Rb <sub>2</sub> S  | _____ |
| 5) MgI <sub>2</sub>  | _____ | 10) Li <sub>2</sub> O | _____ |

#### Part II: Compounds with polyatomic ions

Write the chemical formulas:

Name the following compounds:

- |                            |       |   |       |
|----------------------------|-------|---|-------|
| 11) sodium thiocyanate     | _____ | 21) Mg <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> | _____ |
| 12) magnesium sulfite      | _____ | 22) Al(SCN) <sub>3</sub>                            | _____ |
| 13) ammonium hydroxide     | _____ | 23) KC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>   | _____ |
| 14) calcium nitrate        | _____ | 24) BeCr <sub>2</sub> O <sub>7</sub>                | _____ |
| 15) rubidium chlorite      | _____ | 25) Na <sub>2</sub> CO <sub>3</sub>                 | _____ |
| 16) magnesium oxalate      | _____ | 26) NaHSO <sub>3</sub>                              | _____ |
| 17) lithium nitrite        | _____ | 27) AlPO <sub>3</sub>                               | _____ |
| 18) strontium permanganate | _____ | 28) HClO <sub>2</sub>                               | _____ |
| 19) rubidium phosphite     | _____ | 29) Ba(MnO <sub>4</sub> ) <sub>2</sub>              | _____ |
| 20) ammonium sulfate       | _____ | 30) Sr(CN) <sub>2</sub>                             | _____ |

**Part III: Compounds with transition metals**

Write the chemical formulas:

Name the following compounds:

31) zinc (II) chloride \_\_\_\_\_

41)  $\text{CdF}_2$  \_\_\_\_\_

32) vanadium(II) phosphite \_\_\_\_\_

42)  $\text{CdF}_3$  \_\_\_\_\_

33) vanadium(III) phosphite \_\_\_\_\_

43)  $\text{PtCl}_4$  \_\_\_\_\_

34) manganese (II) fluoride \_\_\_\_\_

44)  $\text{PtCl}_2$  \_\_\_\_\_

35) tin (IV) acetate \_\_\_\_\_

45)  $\text{Fe}_2\text{O}_3$  \_\_\_\_\_

36) tin (II) hypochlorite \_\_\_\_\_

46)  $\text{FeO}$  \_\_\_\_\_

37) copper (II) borohydride \_\_\_\_\_

47)  $\text{VCl}_3$  \_\_\_\_\_

38) nickel(II) oxalate \_\_\_\_\_

48)  $\text{VCl}_5$  \_\_\_\_\_

39) gold (III) sulfite \_\_\_\_\_

49)  $\text{V}_2\text{O}_3$  \_\_\_\_\_

40) lead (II) permanganate \_\_\_\_\_

50)  $\text{VC}$  \_\_\_\_\_

**Part IV: Covalent Compounds**

Name the following compounds:

Write the chemical formulas:

51)  $\text{OF}_2$  \_\_\_\_\_

56) sulfur dioxide \_\_\_\_\_

52)  $\text{N}_2\text{S}_3$  \_\_\_\_\_

57) diphosphorus pentoxide \_\_\_\_\_

53)  $\text{SF}_6$  \_\_\_\_\_

58) oxygen dichloride \_\_\_\_\_

54)  $\text{PCl}_3$  \_\_\_\_\_

59) dinitrogen hexabromide \_\_\_\_\_

55)  $\text{P}_2\text{O}_5$  \_\_\_\_\_

60) carbon tetrachloride \_\_\_\_\_