

Name \_\_\_\_\_

Period \_\_\_\_\_

### Acid Base Nomenclature

1) Name the following binary acids:

a) HF \_\_\_\_\_

d) HI \_\_\_\_\_

b) HCl \_\_\_\_\_

e) H<sub>2</sub>S \_\_\_\_\_

c) HBr \_\_\_\_\_

2) What are the names of the following acids containing the element bromine:

a) HBr \_\_\_\_\_

d) HBrO<sub>3</sub> \_\_\_\_\_

b) HBrO \_\_\_\_\_

e) HBrO<sub>4</sub> \_\_\_\_\_

c) HBrO<sub>2</sub> \_\_\_\_\_

3) What are the formulas of the following acids containing the element chlorine:

a) Hydrochloric acid \_\_\_\_\_

b) Hypochlorous acid \_\_\_\_\_

c) Chlorous acid \_\_\_\_\_

d) Chloric acid \_\_\_\_\_

e) Perchloric acid \_\_\_\_\_

4) What are the names of the following acids containing the element fluorine:

a) HF \_\_\_\_\_

d) HFO<sub>3</sub> \_\_\_\_\_

b) HFO \_\_\_\_\_

e) HFO<sub>4</sub> \_\_\_\_\_

c) HFO<sub>2</sub> \_\_\_\_\_

6) What are the names of the following oxy acids:

a)  $\text{H}_2\text{SO}_4$  \_\_\_\_\_ b)  $\text{H}_2\text{SO}_3$  \_\_\_\_\_

c)  $\text{HNO}_3$  \_\_\_\_\_ d)  $\text{HNO}_2$  \_\_\_\_\_

e)  $\text{H}_3\text{PO}_4$  \_\_\_\_\_ f)  $\text{H}_3\text{PO}_3$  \_\_\_\_\_

g)  $\text{H}_2\text{CO}_3$  \_\_\_\_\_ h)  $\text{H}_2\text{CO}_2$  \_\_\_\_\_

i)  $\text{H}_2\text{CrO}_4$  \_\_\_\_\_ j)  $\text{H}_2\text{Cr}_2\text{O}_7$  \_\_\_\_\_

k)  $\text{HC}_2\text{H}_3\text{O}_2$  \_\_\_\_\_ l)  $\text{H}_2\text{C}_2\text{O}_4$  \_\_\_\_\_

7) What are the names of the following bases:

a)  $\text{NaOH}$  \_\_\_\_\_ f)  $\text{Mg}(\text{OH})_2$  \_\_\_\_\_

b)  $\text{KOH}$  \_\_\_\_\_ g)  $\text{Ba}(\text{OH})_2$  \_\_\_\_\_

c)  $\text{LiOH}$  \_\_\_\_\_ h)  $\text{Ca}(\text{OH})_2$  \_\_\_\_\_

d)  $\text{AgOH}$  \_\_\_\_\_ i)  $\text{Be}(\text{OH})_2$  \_\_\_\_\_

e)  $\text{CuOH}$  \_\_\_\_\_ j)  $\text{Fe}(\text{OH})_2$  \_\_\_\_\_

8) Name the following species that are very common in acid base reactions:

a)  $\text{H}_3\text{O}^+$  \_\_\_\_\_

b)  $\text{OH}^-$  \_\_\_\_\_

c)  $\text{NH}_4^+$  \_\_\_\_\_

d)  $\text{NH}_3$  \_\_\_\_\_

e)  $\text{H}^+$  \_\_\_\_\_